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		C				٥,
	< ml					
	6) There are A) 0	unpaired elec B) 1			F) 4	6)

The principal qua	ntum number for th	e outermost electron	s in a Br atom in the	ground state is	7)
A) 4	B) 1	C) 2	D) 5	E) 3	
The largest princi	pal quantum numbe	er in the ground state	electron configurati	on of barium is	8)
A) 1	B) 2	C) 4	D) 5	E) 6	
What is the correc	et ground-state elect	ron configuration for	: molybdenum?		9)
A) [Kr]5s14d5					
B) [Kr]5s14d10)				
C) [Kr]5s24d9					
D) [Kr]5s24d4					
E) [Kr]5s24d5					
The complete elec	tron configuration o	of gallium, element 3	1, is		10)
A) 1s22s22p10	$3s23p^{10}4s^23d^3$				
B) 1s42s42p83s	$_{\rm s43p}8_{\rm 4s}3$				
C) 1s42s42p63s	*				
•	s23p63d ¹⁰ 4s ² 4p ¹				
E) 1s42s42p10g	• •				
L) 1312312p100	53 5p				
In a ground-state	manganese atoms, t	hesubsh	ell is partially filled.		11)
A) 4s	B) 4p	C) 3s	D) 4d	E) 3d	, <u> </u>
	of an elec	nty Principle, it is in	ipossible to know pr	ecisely bout the	12)
	following is <u>not</u> a va	lid value for the mag	gnetic quantum num	ber of an electron	13)
in a 5d subshell? A) 1	B) 2	C) 0	D) 3	E) -1	
A) 1	D) 2	C) 0	D) 3	E) -1	
An electron canno	ot have the quantum	numbers n =	l =	, mį =	14)
	D) 2 0 0	C) 2 1 1	D) 0 1 1	T\ 1	
A) 3, 2, 1	B) 2, 0, 0	C) 3, 1, -1	D) 2, 1, -1	E) 1, 1, 1	
In a py orbital, the	e subscript x denotes	s the of tl	ne electron.		15)
A) spin of the (B) probability C) size of the (electrons of the shell		ic ciccion.		10,